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Problem : Given that the free energy of formation of liquid water is -237 kJ / mol , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate ΔG for the reaction, which is $-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$. Knowing that $\Delta G = -nFE$ and $n = 4$, we calculate the potential is -1.23 V .

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Thermodynamic Properties

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Answers For Thermodynamics Problems Answer for Problem # 1 Since the containers are insulated, no heat transfer occurs between the gas and the external environment, and since the gas expands freely into container B there is no resistance "pushing" against it, which means no work is done on

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the gas as it expands.

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contents: thermodynamics . chapter 01: thermodynamic properties and state of pure substances. chapter 02: work and heat. chapter 03: energy and the first law of thermodynamics. chapter 04: entropy and the second law of thermodynamics. chapter 05: irreversibility and availability

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