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amounts of products formed by a chemical reaction is called stoichiometry.

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Stoichiometry Practice Worksheet Solve the following stoichiometry grams-grams problems: Using the following equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$ How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid? Using the following equation: $4 \text{LiNO}_3 \rightarrow 2 \text{Li}_2\text{SO}_4$

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Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that

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The ratio of C to A is: $C:A = 3:2$ $C: A = 3: 2$. Hence, the expected yield of C (based on the quantity of A available) is: $3 \times 4 = 6$ units $3 \times 4 = 6$ units . Based on the quantity of B available ...

Let's just consider the concept of stoichiometry without ...

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