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By mixing sodium hydroxide,  $\text{NaOH (aq)}$ , with acetic acid,  $\text{HC}_2\text{H}_3\text{O}_2 \text{ (aq)}$ , no reaction precipitate is observed due to the formation of  $\text{NaC}_2\text{H}_3\text{O}_2 \text{ (aq)}$  which is soluble according to the solubility rule. The reaction can be written as  $\text{NaOH (aq)} + \text{HC}_2\text{H}_3\text{O}_2 \text{ (aq)} \rightarrow \text{NaC}_2\text{H}_3\text{O}_2 \text{ (aq)} + \text{H}_2\text{O (l)}$  (Nothing).

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