

Chemistry Practicals For Class 11 Salt Analysis

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Chemistry Practicals For Class 11

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CBSE Class 11 Chemistry Lab Manual. Determine the strength of the given sodium hydroxide solution. You are provided with oxalic acid solution. Determine the strength and molarity of the given solution of hydrochloric acid. Given is 0.05 M solution. crystals per litre.

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CBSE Chemistry Lab Manual for Class 11 The laboratory is important for making the study complete, especially for a subject like Science and Chemistry. CBSE has included the practicals in secondary class intending to make students familiarised with the basic tools and techniques used in the labs.

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Class 11 Chemistry is a quite complex subject which demands a lot of practice time from the students to master. Moreover, if the student aims to score maximum marks in Class 11 Chemistry Exam then, he/she is required to put in continuous efforts and rigorous practice as this subject has the tendency to challenge students at every step.

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Chemistry Text Book "Lab Manual" ebook for Class for 11 ...

Next we add up $11.2 \text{ mL} + 10.9 \text{ mL} + 11.4 \text{ mL} =$ divide by 3 = 11.16666667 mL *Note we don't have to use 8.9 because it's not a close number to the others Next we change 11.16 mL into L $11.16 \times 1\text{L } 1000 \text{ mL} = .0111\text{L}$ Next we take our concentration and change that into Moles of HI $0.330 \text{ M} \times .011 \text{ L} = .00363 \times 1 = 0.00363 \text{ mol}$

CHEMISTRY 11: TITRATIONS

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