

Chapter Section 12 1 Liquids And Solids Answer Key

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Chapter 12.1 : Liquids
1. Liquids
Chapter 12.1
Objectives
Describe the motion of particles in liquids and the properties of...
2. Liquids
Examples:
Oceans, lakes, rivers
Hard to believe - liquids are the least common state in the...
3. Properties of Liquids and ...

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LIQUIDS AND SOLIDS 363
SECTION 12-1 OBJECTIVES
Describe the motion of particles in liquids and the properties of liquids according to the kinetic-molecular theory. Discuss the process by which liquids can change into a gas. Define vaporization. Discuss the process by which liquids can change into a solid. Define freezing.
Module 1:States of Matter/Classes

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Section 12.1 Gases
Section 12.2 Forces of Attraction
Section 12.3 Liquids and Solids
Section 12.4 Phase Changes
Exit CHAPTER States of Matter 12
Click a hyperlink to view the corresponding slides.
•Use the kinetic-molecular theory to explain the

Chemistry: Matter and Change

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Holt et al. Terms in this set (41) fluid. substance that can flow and take the shape of its container, such as a liquid or a gas.

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The amount of energy requires to stretch or increase the surface of a liquid by a unit area. When a molecule within a liquid is pulled in all directions by the intermolecular forces between it and the other molecules that surround it. No net pull in any one direction. When a molecule at the surface of the liquid is similarly pulled down and to the sides by neighboring molecules, but there is no upward pull to balance the downward or inward pull (into the bulk of the liquid)

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CHAPTER 12 REVIEW Solutions SECTION 2 SHORT ANSWER

Answer the following questions in the space provided. 1. The following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas in the liquid.

CHAPTER 12 REVIEW Solutions

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CHAPTER 12 REVIEW Liquids and Solids SECTION 12-1 SHORT ANSWER Answer the following questions in the space provided.

1. List the following attractive forces in order of increasing strength: (a) hydrogen bonding (c) ionic bonding (b) London dispersion forces (d) dipole-dipole forces 2. All of the following statements about liquids and gases are true except.

12 Liquids and Solids

Chapter 12: Intermolecular Forces: Liquids and Solids 12-1

Intermolecular Forces At high pressures and low temperatures,

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intermolecular forces cause gas behaviour to depart from ideality. When these forces are sufficiently strong compared with the thermal energy, a gas condenses to a liquid. This section will examine the types of intermolecular forces known collectively as van der Waals forces.

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12 Liquids and Solids - colonial.net. Liquids and Solids SECTION 12-3 SHORT ANSWER Answer the following questions in the space provided. 1. Consider the following system at equilibrium: Filesize: 540 KB; Language: English; Published: June 26, 2016; Viewed: 1,853 times

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Chapter 12.1
Objectives
Describe the motion of particles in liquids and the properties of liquids according to the kinetic-molecular theory.
Discuss the process by which liquids can change into a gas.

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