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1 . 5.3 $\times \times 10^3$ C 3 . (a) reduction; (b) oxidation; (c) oxidation; (d) reduction 5 . (a) $F_2 + Ca_2F - + Ca_2+$; $F_2 + Ca_2F - + Ca_2+$; (b) Considerations include: cost of the materials used in the battery, toxicity of the various components (what constitutes proper disposal), should it be a primary or secondary battery, energy requirements (the "size" of the battery/how long ...

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CHAPTER 17 REVIEW Reaction Kinetics MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. The reaction for the decomposition of hydrogen peroxide is $2H_2O_2(aq) \rightarrow 2H_2O(l) + O_2(g)$. List three ways to speed up the rate of decomposition. For each one, briefly explain why it is effective, based on collision theory.

17 Reaction Kinetics - David Brearley High School

CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS 17.1 If the rate of the forward reaction exceeds the rate of reverse reaction, products are formed faster than they are consumed. The change in reaction conditions results in more products and less reactants.