

## Chapter 13 States Of Matter Chemistry Test Answers

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### Chapter 13 States Of Matter

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### States of Matter (chapter 13) Flashcards | Quizlet

You are already familiar with the three common states of matter: solid, liquid, and gas. Solid objects litter the room around you. For example, you can easily recognize the shape of your desk; you know that your backpack cannot hold seven textbooks. You encounter liquids throughout the day as yo u

### Chapter 13: States of Matter

Chapter 13 States of Matter137 SECTION 13.1 THE NATURE OF

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**GASES (pages 385–389)** This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386) 1.

## **Name Date Class STATES OF MATTER 13**

There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These three states are quite different. The main difference is in their particles.

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Chapter 13 States of Matter. STUDY. PLAY. Describes the behavior of gases in terms of particles in motion-Kinetic-Molecular Theory. Gas particles are small and separated from each other by empty space-Particle size. Gas particles in constant, random motion-Particle motion. Move in straight lines-

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Title: Chapter 13 States of Matter 1 Chapter 13States of Matter 2 Kinetic Theory as Applied to Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

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Chemistry (12th Edition) answers to Chapter 13 - States of Matter - 13.1 The Nature of Gases - 13.1 Lesson Check - Page 424 8 including work step by step written by community

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## **Chemistry (12th Edition) Chapter 13 - States of Matter ...**

all matter consists of tiny particles that are constantly in motion  
What are the three assumptions of the kinetic theory as it applies to gases? -The particles in a gas are considered to be small, hard spheres with an insignificant volume. -The motion of the particles in a gas are rapid, constant, and random.

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Chapter 13 - States of Matter. 13.1 The Nature of Gases - Chemistry & You; 13.1 The Nature of Gases - Sample Problem 13.1; 13.1 The Nature of Gases - 13.1 Lesson Check; 13.2 The Nature of Liquids - Chemistry & You; 13.2 The Nature of Liquids - 13.2 Lesson Check; 13.3 The Nature of Solids - Chemistry & You; 13.3 The Nature of Solids - 13.3 Lesson Check; 13.4 Changes of State - Chemistry & You

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Chapter 13 - States of Matter - 13 Assessment - Page 444: 58  
Answer It is called a dynamic equilibrium because vaporization and condensation are still occurring at equal rates even though there is no net change in the number of particles in the liquid or vapor.

## **Chemistry (12th Edition) Chapter 13 - States of Matter ...**

Title: Chapter 13: States of Matter 1 Chapter 13 States of Matter. Kinetic-Molecular Theory Explains the motions and behavior of a gas. The theory has three components ; 1. Particle Size Gas particles are small relative to the space around the particles. This means that there is no significant attraction or repulsion between gas particles. NO transfer of